

PANFULOV, A.

Vybor napravleniya Dono-Volzhskogo soedineniya. [The choice of direction for  
Volga-Don junction]. (Vodnyi transport, 1935, no. 12, p. 14-16, map).  
DLC: HE561.R8

SO: SOVIET TRANSPORTATION AND COMMUNICATIONS, A BIBLIOGRAPHY, Library of Congress  
Reference Department, Washington, 1952, Unclassified.

ZELENSKIY, A.F., PANPULOV, M.S.

Blood - Pressure

Method of measuring arterial pressure of children. *Pediatrila* no. 2, March-April 1952

9. Monthly List of Russian Accessions, Library of Congress, August ~~1953~~<sub>1952</sub>, Uncl.

PANPULOV, M.S., kandidat meditsinskikh nauk. (Moskva)

A self-supporting uterine cannula. Akush. i gin. no. 6:58-60 H-D  
'55 (MIRA 9:6)

(UTERUS, radiography  
contrast media infusion, use of uterine cannula)  
(ROENTGENOGRAPHY, appar. and instruments  
uterine cannula for infusion of contrast media)

PANPURIN, V.N.

Vertical jig for slime water clarification. Sbor.DonUGI no.22:  
129-136 '61. (MIRA 15:6)  
(Separators (Machinery)) (Hydraulic mining)

PANFURINA, I. V. inzhener.

Suggestions for increasing the efficiency of the Rostov Flour Mill  
no. 3. Muk. -elev.prom.22 no 11:24-25 N '56. (MLRA 10:1)

1. Rostovskaya mel'nikarskaya (Rostov-on-Don Flour mills)

STRIMBAN, Yu., inzh.; PANPURINA, L., inzh.

For effective utilization and improvement of MUKZ-35 feed mills.  
Muk.-elev. prom. 25 no.4:17 Ap '59. (MIRA 13:1)

- 1.Sverdlovskoye upravleniye khleboproduktov (for Strimban).
- 2.Rostovskoye upravleniye khleboproduktov (for Panpurina).  
(Feed mills)

PANPURINA, L.

At the Sal'sk Mills. Muk.-elev. prom. 24 no.12:25 D '58.  
(MIRA 12:1)

1. Rostovskoye upravleniye khleboproduktov.  
(Sal'sk--Flour mills)

157 AND 158 CRYSTAL		PROCESSES AND PROPERTIES INDEX		159 AND 160 CRYSTAL	
<p>2-1</p> <p>RELATIVE PERMEABILITY OF SOME OF THE POLYMERIZATION PRODUCTS. I. M. ...            The permeability of the polymerization products between N. ...            and ... was determined by the method of ...            The results of the measurements are given in the table ...            The permeability of the polymerization products for the gases ...            of the same polymerization products is a function of their nature ...            The permeability of the polymerization products of the system ... with ...            is a function of the nature of the system ...            The permeability of the polymerization products for all types of ...            F. O. H.</p>					
<p>ADD-15A METALLURGICAL LITERATURE CLASSIFICATION</p>					
157 AND 158 CRYSTAL		159 AND 160 CRYSTAL		157 AND 158 CRYSTAL	
157 AND 158 CRYSTAL		159 AND 160 CRYSTAL		157 AND 158 CRYSTAL	

GUZINA, Djoko; MILOSAVLJEVIC, Aleksije; PANSEVIC, Mladomir; BUGARSKI,  
Miodrag.

Studies on the blood, erythrocytes, plasma volume and survival  
of erythrocytes in splenomegalic forms of liver cirrhosis.  
Srpski an. celok. lek. 92 no 9:869-877 S'64.

1. Onkološki institut Medicinskog fakulteta Univerziteta u  
Beogradu (Direktor: prof. dr Marija Visnjic-Frajnd).

7

CA

Aluminum-beryllium wrought alloys. C. B. Bessie. *Aluminum* 10, 0 (1935). Several Al-Be alloys were prepared and investigated. Micrographic and tensile data on specimens with different heat-treatment indicate a superior composition: (1) Al 4.5-5, Be 0.3-1.1%, and the test Cu, (2) Al 0.5-7, Be 0.3-0.4%, and the test Cu. Ni or Ag may be added. Beryllium may be decreased if Al is increased; without adversely affecting the phys. properties; this increases resistance against corrosion and tarnishing. Working of these alloys does not present any exceptional difficulties. Frederick C. Nachod.

PANSENKO, Viktor Andreyevich; BELOV, V.S., red. izd-va; PRONINA,  
N.D., tekhn. red.; OVSEYENKO, V.G., tekhn. red.

[Filter protection of electric equipment in mines] Fil'trovaia  
zashchita shakhtnogo elektrooborudovaniia. Moskva, Gosgortekh-  
izdat, 1962. 26 p. (MIRA 15:4)  
(Mining machinery--Electric driving) (Electric filters)

PANSEVICH, V. I.

Pansevich, V. I. "The scientific work of N. A. Prilezhnyev", (Summary of the paper),  
Soobsach. o nauch. rabotakh i izn. Vsesoyuz. Nauch. Sov. in. Kemiye, 1954, Issue 1,  
p. 11-13, - Bibliog. items.

SO: U-361, 1 April 55. (Let's is 'Eurasia' State, N. 11, 1955)





PANSEVICH-KOLYADA, V. I.

Pansevich-Kolyada, V. I. "The scientific work of Mr. A. Prilezmyev", (Chemist), Izvestiya Akad. nauk BSSR, 1948, No. 6, p. 194-195, - Bibliog: 10 items.

SO: U-3261, 10 April 53, (Letovis 'Zhurnal 'nykh Statey, No. 11, 1949.

PANSEVICH-KOLYADA, V. I.

Pansevich-Kolyada, V. I. - "The investigation of allene hydrocarbons", 1; On obtaining methyl phenyl allene", (Report), Soobshch. o nauch. rabotakh chlenov Vsesoyuz. khim. o-va im. Mendeleyeva, 1949, Issue 2, p. 13-14.

SO: U-4630, 16 Sept. 53 (Ietopis 'Zhurnal 'nykh Statey, No. 23, 1949).

PANSEVICH-Kolyada, V. I.

32346

Isslyedovaniye Allyenovy kh uglyevodorodov. Prisoyedineyniye Broma K Dimyetilallyenu. Soobshch. o nauch Rabotakh Chlyenov Vsesoyuz Khim. O-va im. Myendyelyeyeva, 1949 Vyp. 2, s. 51-52.

SO: Letopis' Zhurnal'nykh Statey, Vol. 44, Moskva, 1949

10

CA

Alkene hydrocarbons. I. Preparation of methylphenylallene. V. I. Pansevich-Kolyada. *Zhur. Obshchei Khim.* (J. Gen. Chem.) 21, 494-8 (1951). Bromination of MeEtC(OH)Ph yields 3-bromo-2-phenyl-2-butene, b. 101-4°; this (220 g.) heated 10-15 min. in 600 g. 20% EtOH-KOH gave 122 g. KBr and 144 g. liquid products, b. 42-above 100°; repeated distn. gave an unstated amt. of  $C_{11}H_{10}O$ , b. 90.5-1.0°,  $d_4^{20}$  0.8900,  $n_D^{20}$  1.6382, contg. some Cl deriva. Oxidation with 1%  $KMnO_4$  gave AcPh, HCO<sub>2</sub>H, HCO<sub>2</sub>Et,  $EtOCH_2CO_2H$ , possibly  $HCOC(O)H$ , and  $MePhC(O)H$ .  $EtOCH_2CO_2H$ , b. 111.5-12.5°,  $d_4^{20}$  1.0000,  $n_D^{20}$  1.5170. *Semicarbazone*, m. 162-3°. Hence, the original reaction with KOH-EtOH apparently yielded a *gem-methylphenyl allene*, which under the reaction conditions added EtOH, yielding 2-phenyl-6-ethoxy-2-buten-1-ol, which is the probable structure of the  $C_{11}H_{10}O$  described above. II. Addition of bromine to unsymmetrical dimethylallene. *Ibid.* 498-501. Refluxing  $Me_2C=CHBr$  with alc. KOH gave a mono-Br compd., b. 117-21°,  $d_4^{20}$  1.3073,  $n_D^{20}$  1.4042, apparently  $Me_2C=CHBr$ ; this heated 12 hrs. in a sealed tube with alc. KOH to 140-50° gave 3-methyl-1,2-butadiene, b. 40.5-2.5°. This (3 g.) treated under ice cooling with 7 g. Br gave a range of products, including 2.5 g. 1,2-dibromo-3-methyl-2-buten-1-ol, b. 111-12°,  $d_4^{20}$  1.7702,  $n_D^{20}$  1.5479, oxidized with 1%  $KMnO_4$  to AcH,  $Me_2CO$ , HCO<sub>2</sub>H, and  $BrCH_2CO_2H$ . G. M. Kosolapoff



PANSEVICH-KOLYADA, V.I.; TIMOSHEK, L.I.

Properties of 2,3-epoxy-2-methyl-4-phenyl-4-pentanol. Zhur.. Obshchey  
Khim. 22,1392-6 '52. (MLRA 5:8)  
(CA 47 no.15:7436 '53)

PANSEVICH-KOLYADA, V. I.

Oxides

Obtaining and properties of 1, 1-diphenyl-2-propylene oxide., Zhur. ob. khim., 22, no. 1, 1952.  
Laboratoriya Organicheskoy Khimii Akademii Nauk  
Belorusskoy SSR

SO: Monthly List of Russian Accessions, Library of Congress, May 195<sup>6</sup><sub>3</sub>, Uncl.

PANSEVICH-KOLYADA, V. I.; TIMOSHEK, L. I.

Oxides

Properties of the oxide of 2-methyl-4-phenylpentene-2-ol-4. Zhur.ob.khim. 22 no.8, 1952.

Monthly List of Russian Accessions, Library of Congress, November 1952. Unclassified.

# U.S.S.R.

Oxidation of butadiene hydrocarbons of unsymmetrical structure by  $\text{AcOH}$  and  $\text{AcOH}$  peroxide. V. I. Polterich, Kolyada, T. S. Strigalova, and Z. N. 1418-27 (1953). *Chem. Zh.* 2, 1418-27 (1953). Methyl oxide with  $\text{EtMgBr}$  gave 2,4-dimethyl-2-hexen-1-ol, b.p. 63-4°,  $n_D^{20}$  1.4445,  $d_4^{20}$  0.8493, which dehydrated by boiling 5-6 hrs. to 2,4-dimethyl-2,4-hexadiene, b.p. 115-17°,  $n_D^{20}$  1.4336,  $d_4^{20}$  0.7629 (dehydration with oxalic acid yields much dimerized product, b.p. 113-16°,  $n_D^{20}$  1.4860,  $d_4^{20}$  0.8061). The diene (50 g.) in 150 ml. dry  $\text{Et}_2\text{O}$  treated with  $\text{AcOH}$  (99.79% pure; amt. was based on 2 atoms of O per mole of diene) at 20-5° and kept 6 days gave 2 products: 2,4-dimethyl-2,4-hexadiene dioxide, b.p. 60-2°,  $n_D^{20}$  1.4220,  $d_4^{20}$  0.9629, and  $\text{Me}_2\text{C}=\text{CH}(\text{CMe}(\text{OAc})\text{CH}(\text{OH})\text{Me})_2$ , b.p. 100-4°,  $n_D^{20}$  1.4469,  $d_4^{20}$  1.0380. The dioxide treated with  $\text{AcOH}$  at 100° gave the latter substance, b.p. 100-5°, and much tar. Oxidation of the ethylene oxide deriv. with  $\text{Pb}(\text{OAc})_2$  gave  $\text{AcH}$  as the sole identifiable product. Treatment of the ethylene oxide deriv. with  $\text{Ac}_2\text{O}$  10-12 hrs. at reflux gave a diacetate with preservation of the oxide ring; the product,  $\text{C}_{10}\text{H}_{18}\text{O}_4$ , b.p. 104-5°,  $n_D^{20}$  1.4405,  $d_4^{20}$  1.0690,  $n_D^{25}$  1.4351. Refluxing 2,4-dimethyl-2-octen-1-ol gave 2,4-dimethyl-2,4-octadiene, b.p. 161-3°,  $n_D^{20}$  1.4495,  $d_4^{20}$  0.8254. This oxidized as above with  $\text{AcOH}$  in 11 days gave the dioxide, b.p. 58-64°,  $n_D^{20}$  1.4330,  $d_4^{20}$  0.8390. The monoacetates were  $\text{C}_{12}\text{H}_{20}\text{O}_4$  (the structure of this and other monoacetates were not firmly established), b.p. 126-30°,  $n_D^{20}$  1.4500,  $d_4^{20}$  1.0155; the latter with  $\text{Pb}(\text{OAc})_2$  gave  $\text{PrClO}$ , while treatment with  $\text{Ac}_2\text{O}$  as above gave a diacetate,  $\text{C}_{14}\text{H}_{26}\text{O}_6$ , b.p. 121-3°,  $n_D^{20}$  1.4433,  $d_4^{20}$  1.0351. Treating the monoacetate with  $\text{MeOH}$  contg. a little  $\text{H}_2\text{SO}_4$  at 35° gave after 3 hrs. a Me

ether of a keto glycol,  $C_{17}H_{26}O_3$ , b<sub>p</sub> 103-104°, n<sub>D</sub><sup>20</sup> 1.4661, d<sub>4</sub><sup>20</sup> 1.0112; 2,4-dinitrophenylhydrazones, m: 131-132°. Iso-Am-MgBr and methyl oxide gave 2,4,7-trimethyl-2-octen-1-ol, b<sub>p</sub> 65°, n<sub>D</sub><sup>20</sup> 1.4488, d<sub>4</sub><sup>20</sup> 0.8427,  $\gamma_{20}^{20}$  20.24, which heated to 100° with MgSO<sub>4</sub> gave 2,4,7-trimethyl-2,4-octadiene, b<sub>p</sub> 63-64°, n<sub>D</sub><sup>20</sup> 1.4480, d<sub>4</sub><sup>20</sup> 0.7731. This with AcOH as above gave in 7 days the dioxide, b<sub>p</sub> 63-68°, n<sub>D</sub><sup>20</sup> 1.4827, d<sub>4</sub><sup>20</sup> 0.8617, and the oxide glycol monoacetate,  $C_{17}H_{30}O_4$ , b<sub>p</sub> 119-20°, n<sub>D</sub><sup>20</sup> 1.4400, d<sub>4</sub><sup>20</sup> 0.9982. Oxidation of the latter with Pb(OAc)<sub>2</sub> gave iso-BuCHO (2,4-dinitrophenylhydrazone, m: 152-3°), and treatment with  $As_2O_3$  gave the monoacid diacetate,  $C_{17}H_{30}O_6$ , b<sub>p</sub> 126-36°, n<sub>D</sub><sup>20</sup> 1.5447, d<sub>4</sub><sup>20</sup> 1.0287. Heating 2-phenyl-2-butanol with (CO<sub>2</sub>H)<sub>2</sub> 6 hrs. at 100° gave 2-phenyl-2-butene, b<sub>p</sub> 74.5-75°, n<sub>D</sub><sup>20</sup> 1.5340, d<sub>4</sub><sup>20</sup> 0.9140. This with AcOH as above gave in 2 days the oxide, b<sub>p</sub> 77-80°, n<sub>D</sub><sup>20</sup> 1.5162, d<sub>4</sub><sup>20</sup> 1.0098,  $\gamma_{20}^{20}$  34.89, and PhMeC(OAc)CH(OH)Me (I), b<sub>p</sub> 130-2°, n<sub>D</sub><sup>20</sup> 1.5184, d<sub>4</sub><sup>20</sup> 1.0395,  $\gamma_{20}^{20}$  39.41. Oxidation of the latter with Pb(OAc)<sub>2</sub> gave AcPh. The oxide with MeOH in the presence of H<sub>2</sub>SO<sub>4</sub> gave MePhC(OAc)CH(OH)Me, b<sub>p</sub> 92-4°, n<sub>D</sub><sup>20</sup> 1.5160, d<sub>4</sub><sup>20</sup> 1.0447,  $\gamma_{20}^{20}$  37.22. The oxide with EtOH in the presence of H<sub>2</sub>SO<sub>4</sub> gave the isomerized product MePhCHAc, b<sub>p</sub> 74-8°, n<sub>D</sub><sup>20</sup> 1.5170, d<sub>4</sub><sup>20</sup> 0.9804; semicarbazone, m: 146-9°. I with MeOH in the presence of H<sub>2</sub>SO<sub>4</sub> gave 2-phenyl-2,3-butanediol, b<sub>p</sub> 124-5°, n<sub>D</sub><sup>20</sup> 1.5209, d<sub>4</sub><sup>20</sup> 1.0908.

G. M. Kosolapoff

PANSEVICH-KOLYADA, V.I.

USSR/Chemistry - Oxidation

Card 1/1 Pub. 151 - 8/38

Authors : Pansevich-Kolyada, V. I., and Kureychik, L. A.

Title : Investigation of alcohol oxides (oxidols). Part 2.- Derivation and properties of alpha,beta-alcohol oxides of the aliphatic series

Periodical : Zhur. ob. khim. 24/2, 231-234, Feb 1954

Abstract : The oxidation of 2,4,5-trimethylhexene-2-ol-4 and 2,4-dimethyldecene-2-ol-4 with  $\text{ACH}_2\text{O}_2$  and the characteristics of alpha,beta-alcohol oxides, obtained as result of the oxidation, are described. The oxidation was carried out in absolute ether at 20-25°. Exposure of 2,4,5-trimethyl-oxido-2,3-hexanol-4 to the effects of acetic anhydride resulted in acetylation of the alcohol group without any change to the oxide ring. Exposure of the alpha,beta-alcohol oxides to the effects of anhydrous  $\text{ZnCl}_2$  resulted in the splitting of the alcohol oxides into isobutyric aldehyde and homologous ketones. Eleven references: 9-USSR and 2-German (1860-1952).

Institution : Academy of Sciences Byeloruss-SSR, Institute of Chemistry, Laboratory of Organic Chemistry

Submitted : October 2, 1953

PANSEVICH - SYM#1075  
KOLYADA V.I.

USSR/Chemistry - Synthesis

Card 1/1 : Pub. 151 - 19/37

Authors : Pansevich-Kolyada, V. I., and Ahlova, V. A.

Title : Investigation of alcohol oxides (Oxidols). Part 3.-Derivation and properties of 2-methyl-4-benzyl-oxido-2,3-pentanol-4

Periodical : Zhur. ob. khim. 24/3, 493-498, Mar 1951

Abstract : Investigation of 2-methyl-4-benzyl-oxido-2,3-pentanol-4 showed that the properties of alpha, beta-alcohol oxides are determined not only by the presence and orientation of their functional groups but also by the radicals adjoining the oxidized carbon atoms. Removal of the phenyl radical from the tertiary alcohol group stabilizes the carbon skeleton of the alcohol oxide. The difference between this alcohol oxide and oxides of the aliphatic series, as well as alcohol oxides of 2-methyl-4-phenyl-oxido-2,3-pentanol-4, in which the phenyl radical adjoins directly the carbon atom of the tertiary alcohol group, is explained. Five references: 4-USSR and 1-German (1906-1954).

Institution : Acad. of Sc. Byeloruss-SSR, Institute of Chemistry, Laboratory of Org. Chem.

Submitted : October 29, 1953

PANSEVICH-KOLYADA, V. I.  
USSR/Chemistry

Card 1/1

Authors : Pansevich-Kolyada, V. I.; and Idel'chik, Z. B.

Title : Investigation of alcohol-oxides (oxidols). Part 4.- Properties of alpha-oxides of allyl- and propenyl derivatives of phenol and o-cresol.

Periodical : Zhur. Ob. Khim. 24, Ed. 5, 807 - 814, May 1954

Abstract : Hydrogen peroxide oxidation of allyl- and propenyl derivatives of phenol and o-cresol led to the derivation of alpha-oxides of phenol allyl ether, o-allylphenol, o-allyl-o-cresol, o-propenylphenol and o-propenyl-o-cresol. Allyl phenol oxides do not change under the effect of zinc chloride but when heated with acetic anhydride they acetylize according to the phenol hydroxyl with the formation of oxido-acetates. The conversion of o-propenylphenol oxides into coumarone derivatives confirms the presence of an oxide-enol tautomerism for the alcohol-oxides. Nine references; 2 USSR since 1898.

Institution : Acad. of Sci. Byeloruss-SSR, Institute of Chemistry, Laboratory of Organic Chemistry

Submitted : December 7, 1953

**PANSEVICH-KOLYDA, V.I.**

USSR/Chemistry - Hydrocarbon oxidation

Card 1/1 : Pub. 151 - 27/42

Authors : Pansevich-Kolyada, V. I., and Idelchik, Z. B.

Title : Study of allene hydrocarbons. Part 3.- Oxidation of allene hydrocarbons with acetyl hydrogen peroxide

Periodical : Zhur. ob. khim. 24/49, 1617-1624, Sep 1954

Abstract : The oxidation with acetyl hydroperoxide of three allene hydrocarbons - 2-methylhexadiene-2,3, 2-methyloctadiene-2,3 and 2-methyl-4-phenylpenta- diene-2,3, was investigated. It was found that the oxidation of the allene hydrocarbons with acetyl hydroperoxide takes place in both double bonds. The products obtained from such oxidation are described. The application of the acetyl hydroperoxide oxidation method, as a means of analyzing allene and isomeric acetylene hydrocarbon mixtures, is recommended. Twenty-four references: 17-USSR; 3-USA and 4-German (1872-1954).

Institution : Academy of Sciences Byeloruss-SSR, Institute of Chemistry

Submitted : April 15, 1954

Pansevich-Kolvada, V.I.

Preparation of some esters with allylic position of a double bond. V. I. Pansevich-Kolvada, Z. N. 140' (1955). — To 8.5 g. 2,4-dimethyl-2-hexen-4-ol in 20 ml. MeOH was added 5 ml. 1:5 H<sub>2</sub>SO<sub>4</sub>, yielding after several min. a 2-layered mixt. After diln. the upper layer gave 70% 3,4-dimethyl-4-methoxy-2-hexene, b<sub>p</sub> 40°, d<sub>4</sub> 1.47-8°, n<sub>D</sub> 1.4513, d<sub>20</sub> 0.8192°. Use of EtOH gave 80% Me<sub>2</sub>C:CHCHMe(OEt)Et, b<sub>p</sub> 53.5°, n<sub>D</sub> 1.430°, d<sub>4</sub> 0.8140; allyl alc. gave 53.4% Me<sub>2</sub>C:CHCHMe(OCH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub>)Et, b<sub>p</sub> 53-5°, n<sub>D</sub> 1.4480, d<sub>4</sub> 0.8381. Similarly MeOH and PhCH:CHCH-MeOH gave 91.3% PhCH:CHCHMeOMe, b<sub>p</sub> 114-17°, n<sub>D</sub> 1.5654, d<sub>4</sub> 0.9067.

G. M. Koroladoff

Allen, hydrocarbons. III. Oxidation of allenic hydrocarbons with acetyl hydroperoxide. V. I. Prigodnyy and Z. M. Ilyukhina. J. Gen. Chem. (USSR) 49:1242, 1975 (1975) (Engl. translation). See also 49:1242.

R.M.1

①  
M. Ilyukhina

Panasevich-Kolvaad, V.I.

Oxide compounds. V. Mechanism of cleavage of  $\alpha,\beta$ -alcohol oxides of the aliphatic series under the action of zinc chloride and aqueous solutions of mineral acids. V. I. Panasevich-Kolvaad. *Zhur. Obshchei Khim.* 25, 2001-2002 (1951), 49, 2024; *1834*.  $\alpha,\beta$ -Dipoxy alcs are cleaved by  $ZnCl_2$  or aq. acids to ketones and aldehydes corresponding to the structure of the alc. The probable mechanism of the cleavage is suggested. Oxidation of 2,4-dimethyl-2-penten-4-ol with  $AcO_2H$  in  $Et_2O$  at  $21^\circ$  gave the epoxide, b.  $40-52^\circ$ ,  $n_D^{20}$  1.4165,  $d_4^{20}$  0.9384. This (lit.) with 20 ml.  $H_2O$  and 0.2 g.  $H_2SO_4$  gave in 2 hr. at  $100^\circ$   $Me_2CO$  and  $iso-PrCHO$ , along with some higher-boiling products. b.  $191-211^\circ$ .  $E(MgBr)$  with mesityl oxide gave 2,4-dimethyl-2-hexen-4-ol, b.  $53-4^\circ$ ,  $n_D^{20}$  1.445,  $d_4^{20}$  0.9463, which with  $AcO_2H$  gave the epoxide, b.  $51.5^\circ$ ,  $n_D^{20}$  1.4301,  $d_4^{20}$  0.9288; this heated with  $ZnCl_2$  on a steam bath gave  $iso-PrCHO$  and  $Me_2CO$ ; the same formed on heating with aq.  $H_2SO_4$ , in addn. to a substance  $C_8H_{10}O$ ,  $n_D^{20}$  1.4463,  $d_4^{20}$  0.9525, which forms a 2,4-dinitrophenylhydrazone.  $E(MgBr)$  and mesityl oxide gave 2,4-dimethyl-2-hepten-4-ol, b.  $40-2^\circ$ ,  $n_D^{20}$  1.4453,  $d_4^{20}$  0.9521, which with  $AcO_2H$  gave the epoxide, b.  $52^\circ$ ,  $n_D^{20}$  1.4317,  $d_4^{20}$  0.9164; this heated with  $ZnCl_2$  gave  $iso-PrCHO$  and  $Me_2CO$ ; aq.  $H_2SO_4$  gave the same products, as well as  $C_8H_{10}O$ , b.  $92-4^\circ$ ,  $n_D^{20}$  1.4438,  $d_4^{20}$  0.9363.  $BuMgBr$  and mesityl oxide gave 2,4-dimethyl-2-octen-4-ol, b.  $40^\circ$ ,  $n_D^{20}$  1.4695,  $d_4^{20}$  0.9540, which with  $AcO_2H$  gave the epoxide, b.  $76-7^\circ$ ,  $n_D^{20}$  1.4344,  $d_4^{20}$  0.9059, which with aq.  $H_2SO_4$  gave  $iso-PrCHO$  and  $Me_2CO$ , and some  $Cr$ .

(over)

$M_2O_3$ , b<sub>p</sub> 115-10°, n<sub>D</sub><sup>20</sup> 1.4189, d<sub>4</sub> 0.9393, the same being formed on heating with ZnCl<sub>2</sub>. Heating the oxide with aq. NaOH gave no reaction, and heating in sealed tube with Et<sub>3</sub>NH 6 hrs. at 140° was also ineffective. Iso-AmMgBr and mesityl oxide gave 2,4,7-trimethyl-3-octen-2-ol, b<sub>p</sub> 63°, n<sub>D</sub><sup>20</sup> 1.4458, d<sub>4</sub> 0.8427, which with AcOH gave the epoxide, b<sub>p</sub> 71-2°, n<sub>D</sub><sup>20</sup> 1.4340, d<sub>4</sub> 0.8967, which heated with aq. H<sub>2</sub>SO<sub>4</sub> gave iso-PrCHO and iso-AmCOMe, as well as some C<sub>10</sub>H<sub>18</sub>O<sub>2</sub>, b<sub>p</sub> 100-2°, n<sub>D</sub><sup>20</sup> 1.4489, d<sub>4</sub> 0.9191. G. M. K.

2  
1

*[Handwritten signature]*

Pansevich - Kolyada, V.I.

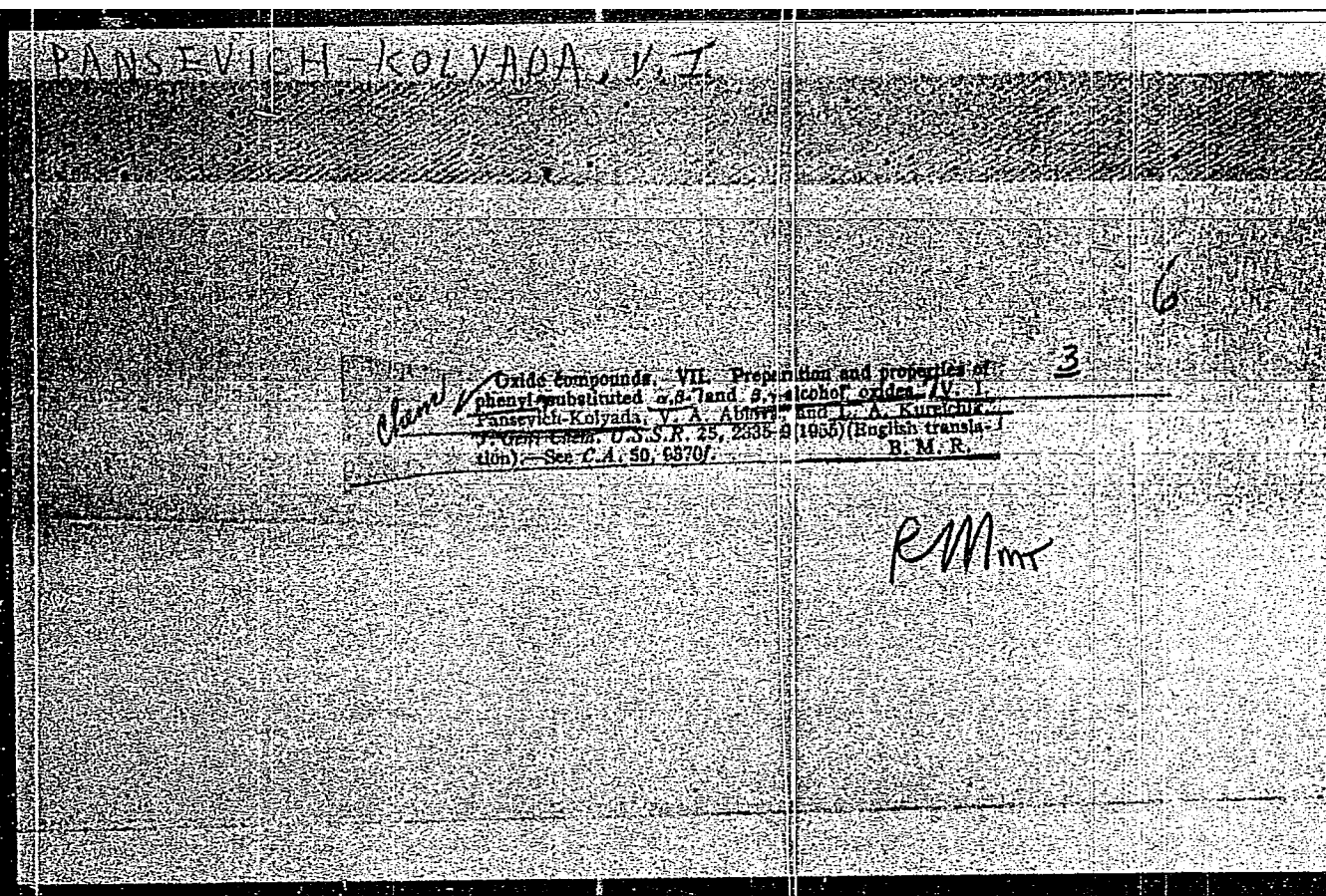
Oxide compounds. VI. Properties of  $\alpha$ -oxides of allyl  
and propenyl derivatives of *m*- and *p*-cresols and gualacol.  
V. I. Pansevich-Kolyada and Z. B. Idel'chuk. *J. Gen.  
Chem. U.S.S.R.* 25, 2177-83(1955)(Engl. translation).—  
See *C.A.* 50, 9370b.

B. M. H.

PAKSEVICH-KOLYADA, V.I.; IDEL'CHIK, Z.B.

Research in the field of oxido compounds. Part 6. Properties of  $\alpha$ -oxides of allyl- and propenyl derivatives of *m*-, *p*-cresols and of guaiacol. Zhur.ob.khim. 25 no.12:2215-2222 N '55.  
(MIRA 9:4)

1. Institut khimii Akademii nauk Belorusskoy SSR.  
(Oxides) (Creosol) (Guaiacol)



**PANSEVICH-KOLYADA, V.I., ABLOVA, V.A., KUREYCHIK, L.A.**

Research in the field of oxido compounds. Part 7. Preparation  
and properties of phenyl substituted  $\alpha,\beta$ - and  $\beta,\gamma$ -alcohol oxides.  
Zhur.ob.khim. 25 no.13:2448-2453 D '55. (MLRA 9:3)

1. Institut khimii Akademii nauk Belorusskoy SSR.  
(Alcohols)

PANSEVICH-KOLYADA, V. I.

*Diene*  
Oxidation of alkadiene hydrocarbons of asymmetric structure by means of acetyl hypoperoxide, II, V. I. Pansevich-Kolyada. *Zhur. Obshch. Khim.* 26, 2181-76 (1950); *U.S.S.R.* 49, 4815f. — Oxidation of unsym. dienes with  $\text{AcO}_2\text{H}$  proceeds by 2 parallel paths: in the first phase of one path there occurs the oxidation of 1 double bond to the monoxide, followed by formation of the dioxide; in the other path the first phase consists of addn. of the  $\text{AcO}_2\text{H}$  across 1,4-positions of the diene, which is followed by oxidation of the dieneol monoacetate to the corresponding monoacetate of an epoxy analog. The selected dienes were oxidized with a calcd. amt. of  $\text{AcO}_2\text{H}$  5-10 days in dry  $\text{Et}_2\text{O}$  with the aim of affecting both double bonds, and after the neutralization and distn. the following dioxides and monoacetates were isolated from the corresponding

dienes:  $\text{O.CMe}_2.\text{CHCMe}_2.\text{CH}_2.\text{O}$  (11 g. from 46 g. diene),  $b_p$  64-7°,  $n_D^{20}$  1.4213,  $d_4^{20}$  0.9608 ( $n_D^{20}$ ,  $d_4^{20}$  were also in-

dicated below);  $\text{O.CMe}_2.\text{CHCMe}_2.\text{CH}_2.\text{O}$  (8 g. from

40 g. diene),  $b_p$  64-6°, 1.4320, 0.9473;  $\text{O.CMe}_2.\text{CH}_2$

$\text{CMe}_2.\text{CMe}_2.\text{O}$  (17 g. from 70 g. diene);  $b_p$  78.5-80°, 1.4325

PANSEVICH-KOLYADA, V. I.

1.0281;  $\text{AcOCH}_2\text{CH}_2\text{CMe}(\text{CHMeOH})_2$  (34.5 g. from 70

g. diene), b<sub>p</sub> 120-3°, 1.4460, 1.0435;  $\text{AcOCH}_2\text{CH}_2\text{CMe}$

$(\text{CHMeOH})_2$  (I) (51.5 g. from 93.5 g. diene), b<sub>p</sub> 126-

8°, 1.4543, 0.9811;  $\text{AcOCH}_2\text{CH}_2\text{CMe}(\text{CHMeOH})_2$ , b<sub>p</sub>

99-101°, 1.4430, 1.0543. I heated with  $\text{Ac}_2\text{O}$ , finally re-

fluxed 7 hrs. gave 2,4-dimethyl-2,5-diacetoxy-3,6-oxydecane,

b<sub>p</sub> 140.5-5°, 1.4481, 1.0000. Reaction of 50 g. 2,4-di-

methyl-2,4-hexadiene with 48 g. 73.8%  $\text{AcOH}$  in 140 ml.

$\text{Et}_2\text{O}$  gave in 2 days at 27-30° 2,4-dimethyl-2,5-oxy-3-hexene,

b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene

underwent some change on heating to 140-150° (140-150°), the results were similar

for oxidation in  $\text{Et}_2\text{O}$  at 9° or in  $\text{AcOH}$ . II treated with

$\text{PhIOAc}$  at 60-65° gave carbonyl compds. which were

unidentified otherwise but that they form 2,4-dimethyl-2,5-

hydroxy-3-hexene, m<sub>p</sub> 124-30° and 136-40°; nonvolatile products

contained 2,4-dimethyl-2,5-diacetoxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.) with

24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave in 24 hrs. 2,4-dimethyl-2,5-

oxy-3-hexene, b<sub>p</sub> 118-21°, 1.4430, 1.0067. 2,4-Dimethyl-2,5-

oxy-3-hexene (30 g.) with 24 g. 80%  $\text{AcOH}$  in 100 ml. gave

in 24 hrs. 2,4-dimethyl-2,5-oxy-3-hexene, b<sub>p</sub> 118-21°;

1.4430, 1.0067. 2,4-Dimethyl-2,5-oxy-3-hexene (30 g.)

PANSEVICH-KOLYADA, V. I.

✓ Action of acetyl hydroperoxide on alkylfuryl alcohols.  
M. M. Azanovskaya and V. I. Pansevich-Kolyada. Dokl. Akad. Nauk S.S.S.R. 111, 1235-1236 (1966). Alkylfurylcarbinols were treated with 90-8%  $\text{AcO}_2\text{H}$  in  $\text{Et}_2\text{O}$  at 20-5° with 1:1 and 1:2 molar proportions of the reactants. With 1:1 mole ratio there were formed 2,3-epoxy-2-furylalkylcarbinols (alkyl group shown): *Pr*, 48%, m. 69.5-71°; *Pr*, 62.7%, m. 67.5-68°; *Bu*, 72.8%, m. 82-3°; *iso-Am*, 30%, m. 60-1.5°. Treatment of the *Bu* compd. with  $\text{ZnCl}_2$  or prolonged storage resulted in decamps. yielding  $\text{BuCHO}$ . When 2 moles of  $\text{AcO}_2\text{H}$  is used for the oxidation only the *Bu* compd. gave a trace of the above described monoepoxy compd. The main bulk of the material from such reactions consisted of mixts. of aldehydes and acids. Thus the *Bu* compds. gave  $\text{BuCHO}$ ,  $\text{HCO}_2\text{H}$ ,  $\text{AcOH}$ , and unidentified acids. The *Et* compd. gave  $\text{EtCHO}$ ,  $\text{HCO}_2\text{H}$ , and  $\text{AcOH}$ , as well as unidentified acids. When the reaction was stopped before completion, appreciable amts. of monoepoxy compds. could be isolated. G. M. Koslanoff

*PANSEVICH-KOLYADA, V. I.*

**AUTHORS:** Azanovskaya, M. M. and Pansevich-Kolyada, V. I. 79-2-24/58

**TITLE:** Alpha-Oxides of Alkyfuryl Alcohols (Alfa-Oksi alkilfurilovykh spirtov)

**PERIODICAL:** Zhurnal Obshchey Khimii, 1957, vol 27, No 2, pp. 384-387 (U.S.S.R)

**ABSTRACT:** Since alcohol oxides of various structure are quite different from each other by their properties, the authros decided to synthesize alcohol oxides containing the furan cycle in the molecule. Oxidation of ethylfuryl, n-propylfuryl, n-butylfuryl and isoamylfuryl alcohols with acetyl hydrogen peroxide in the exact molecular ratios of the alcohol and hydrogen peroxide is described. The products obtained from the oxidation of the alcohols are listed as monoxides of the very same alcohols: 2-(1-hydroxypropyl)-oxido-2,3-furan, 2-(1-hydroxybutyl)-Oxido-2,3 furan, 2(1-hydroxyamyl)-oxido-2,3-furan and 2-(4-methyl-1-hydroxyamyl) oxido-2,3-furan. The physico-chemical properties of the alcohol oxides are described.

Card 1/2

Alpha-Oxides of Alkylfuryl Alcohols

79-2-24/58

There are 11 references, of which 7 are Slavic

ASSOCIATION: Academy of Sciences of Byelorussian-SSR, Institute of Chemistry

PRESENTED BY:

SUBMITTED: March 24, 1956

AVAILABLE: Library of Congress

Card 2/2

*PANSEVICH-KOLYADA V. I.*

79-2-55/64

AUTHOR: Pansevich-Kolyada, V. I.,

TITLE: 2-Phenyl-3-Bromobutene-2 in the Grignard-Reaction. (2-fenil-3-brombuten-2- v reaktsii Grin'yara)

PERIODICAL: Zhurnal Obshchey Khimii, 1958, Vol. 28, Nr 2, pp. 438-441 (USSR)

ABSTRACT: If the above mentioned bromide (I) is added to magnesium in ether the reaction begins immediately and passes energetically to the end whereas only half of the magnesium reacts. A hydrocarbon of the bruttoformula  $C_{20}H_{22}$  (II) is obtained as main product of the reaction i.e. the hydrocarbon skeleton of the bromide is doubled. In order to detect the constitution of (II) it was oxidized according to different methods, and with all of them diphenylacetyl and acetic acid were proved. The compound 3,6-diphenyloctadiene-2,6 (III) is to be considered the main product of the reaction of (I) with magnesium according to this and earlier statements. The production of (III) passes various intermediate stages (tautomers) and side reactions. The dioxide 3,6-diphenyl-dioxydo-2,3-6,7-octane was obtained in the oxidation of (III) by acetylhydroperoxides. This compound is not stable and decomposes in vacuum distillation. It was not possible to obtain derivatives of the dioxide. The specific pro-

Card 1/2

2-Phenyl-3-Bromobutene-2 in the Grignard-Reaction.

79-2-35/64

perties of the obtained compounds (boiling limits of the distillates etc.) as well as the working methods for the reactions are given. There are 10 references, 5 of which are Slavic.

ASSOCIATION: ~~Belorussian~~ Polytechnical Institute (Belorusskiy politekhnicheskiy institut)

SUBMITTED: February 10, 1957

AVAILABLE: Library of Congress

Card 2/2

5.3400

77871

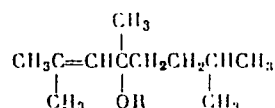
SOV/79-30-2-22/78

AUTHORS: Pansevlch-Kolyada, V. I., Galysheva, T. A.

TITLE: Study of Ethers With an Allylic Double Bond. VI.  
Ethers of the 2,4,7-Trimethyl-2-octen-4-ol

PERIODICAL: Zhurnal obshchey khimii, 1960, Vol 30, Nr 2, pp 460-472  
(USSR)

ABSTRACT: The authors have synthesized 10 new ethers by reacting  
2,4,7-trimethyl-2-octen-4-ol (I) with various saturated  
alcohols (II-XI):



(II) R = CH<sub>3</sub>; (III) R = C<sub>2</sub>H<sub>5</sub>; (IV) R = n-C<sub>3</sub>H<sub>7</sub>; (V) R = n-C<sub>4</sub>H<sub>9</sub>; (VI) R = i-C<sub>4</sub>H<sub>9</sub>; (VII) R =  
= t-C<sub>4</sub>H<sub>9</sub>; (VIII) R = C<sub>5</sub>H<sub>11</sub>; (IX) R = C<sub>6</sub>H<sub>13</sub>; (X) R = CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>; (XI) R = CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>.

Card 1/4

Study of Ethers With an Allylic Double Bond.

77871

VI. Ethers of the 2,4,7-Trimethyl-2-octen-4-ol

SOV/79-30-2-22/78

Etherification went smoothly at room temperature in presence of sulfuric acid in the alcohol mixtures (see Pansevich-Kolyada, V. I., et al., Zhur. obshchey khim., 25, 1481 (1955); 28, 641, 909 (1958)). Constants and yields of the synthesized alcohols are given in Table A. It can be seen that, with increase in molecular weight of R, the yield of the ether is lowered. Etherification of (I) with ethylene glycol or glycerol yields only monoethers. Oxidation of 2,4,7-trimethyl-4-methoxy-2-octene (II) (by reacting for 2 days 850 ml of 3%  $\text{KMnO}_4$  with 10.3 g of ether emulsified in 100 ml of water) yielded 2,4,7-trimethyl-4-methoxy-2-octanol-3-one, which had the following properties: bp 100-102° (3 mm);  $n_D^{20}$  1.4361;  $d_4^{20}$  0.9325. There is 1 table; and 9 references, 8 Soviet, 1 U.S. The U.S. reference is: G. E. Goltz, D. N. Glew, Anal. Chem., 29, 816 (1957).

ASSOCIATION:

Belorussia Polytechnical Institute (Belorusskiy politekhnicheskiy institut)

SUBMITTED:

February 6, 1959

Card 2/4

77871, SOV/79-30-2-22/78

Table A (Caption on Card 3/4)

(1)	(2)	(3) $n_D^{20}$	(4) $d_4^{20}$	(5) MR <sub>D</sub>	(6)	(7)	(8)	(9)	(10)	(11)	(12)	(13)	(14)
C <sub>11</sub> H <sub>20</sub> O (II)	58—59 <sup>0</sup> (2)	1.4380	0.8221	58.73	58.79	78.39	13.27	78.26	13.04	82.4			
C <sub>12</sub> H <sub>20</sub> O (III)	64—65 (2)	1.4390	0.8208	63.45	63.41	78.57	13.38	78.78	13.13	68.0			
C <sub>14</sub> H <sub>20</sub> O (IV)	71—74 (1)	1.4302	0.8234	67.78	68.02	79.51	13.22	79.24	13.20	63.0			
C <sub>15</sub> H <sub>20</sub> O (V)	95—97.5 (4)	1.4395	0.8229	72.54	72.64	79.51	13.24	79.61	13.61	44.7			
C <sub>17</sub> H <sub>20</sub> O (VI)	86—89 (4)	1.4371	0.8171	72.49	72.61	79.97	13.15	79.61	13.27	40.0			
C <sub>17</sub> H <sub>20</sub> O (VII)	102—103.5 (3)	1.4385	0.8179	77.26	77.10	80.21	13.59	80.00	13.33	35.2			
C <sub>18</sub> H <sub>20</sub> O (VIII)	77.5—79 (2)	1.4468	0.8320	67.42	67.56	79.90	12.50	80.00	12.70	66.9			
C <sub>14</sub> H <sub>20</sub> O (IX)	150—152 (3)	1.4902	0.9232	81.46	81.25	83.33	11.21	83.07	10.76	2.9			
C <sub>16</sub> H <sub>20</sub> O (X)	173 (2)	1.4572	0.8703	64.93	66.98	73.20	12.45	72.89	12.15	35.7			
C <sub>17</sub> H <sub>20</sub> O <sub>2</sub> (XI)	194—197 (2)	1.4635	0.9451	70.11	71.07	69.14	11.29	68.88	11.47	52.7			

C<sub>14</sub>H<sub>20</sub>O<sub>2</sub> (XI)  
Card 4/4

PANSEVICH-KOLYADA, V.I.; SHNYP, I.A.

Ethers with the allyl position of a double bond. Part 17: Ethers  
of 2,4-dimethyl-2,6-heptadien-4-ol. Zhur. org. khim. 1 no.1:57-59  
Ja '65. (MIRA 18:5)

1. Belorusskiy politekhnicheskiy institut.

PANSEVICH-KOLYADA, V.I.; GALYSHEVA, T.A.

Oxydo compounds. Part 18: Interaction of monoalkylsubstituted  
glycidols with aniline and piperidine. Zhur. org. khim. 1 no.1:  
172-174 Ja '65. (MIRA 18:5)

1. Belorusskiy politekhnicheskiy institut.

PANSEVICH-KOLYADA, V.I.; BOGUSH, B.K.

Study of ethers with the allyl position of a double bond. Part 8:  
Etherification of some di- and tetra-substituted allyl alcohols with  
primary aliphatic alcohols. Zhur.ob.khim. 33 no.7:2137-2139 J1  
'63. (MIRA 16:8)

1. Belorusskiy politekhnicheskiy institut.  
(Allyl alcohol) (Etherification)

PANSEVICH-KOLYADA, V.I.; BOGUSH, B.K.

Oxido compounds. Part 14: Preparation of some phenyl-substituted ~~alcohol~~-alcohol oxides and their reactions with aniline and o-toluidine. Zhur.ob.khim. 32 no.11: 3552-3556 N '62. (MIRA 15:11)

1. Belorusskiy politekhnicheskii institut.  
(Ethers) (Aniline) (Toluidine)

PANSEVICH-KOLYADA, V.I.

Structure of the bromide  $C_{10}H_{11}Br$ , a product of  
bromination of 2-phenyl-2-butanol and 2-phenyl-2-butene.  
Zhur.ob.khim. 32 no.10:3301-3303 0 '62. (MIRA 15:11)

1. Belorusskiy politekhnicheskiy institut.  
(Butanol) (Butene) (Bromination)

GURINOVICH, I.F.; PANSEVICH-KOLYADA, V.I.

Nature of the hydrogen bond of  $\alpha\beta$ -alcohol oxides and some  
unsaturated tertiary alcohols. Zhur.fiz.khim. 35 no.12:2754-  
2758 D '61. (MIRA 14:12)

1. Akademiya nauk BSSR, Institut fiziki.  
(Hydrogen bonding) (Alcohols)

GUSENITSA, M.I.; ~~PANSEVICH-KOLYADA, V.I.~~

Method for clarifying fruit and berry juice. Sbor.nauch.trud.  
Bel. politekh.inst. no.87,76-78 '59. (MIRA 14:4)  
(Liquids—Clarification)

PANSEVICH KOLYADA, V.I.

Chemistry of 1,2-diene (allene) hydrocarbons. Sbor.nauch.trud.Bel.  
politekh.inst. no.87:3-38 '59. (MIRA 14:4)  
(Propadiene)

OSIPENKO, I.F.; PANSEVICH-KOLYADA, V.I.

Hydroxy compounds. Properties of  $\alpha, \alpha, \gamma, \gamma$ -tetraalkyl substituted  
glycidyl ethers. Sbor. nauch. rab. Inst. fiz.-org. khim. AN BSSR  
no. 7:78-88 '59. (MIRA 14:4)

(Ethers)

PANSEVICH-KOLYADA, V.I.; STREL'TSOV, A.Ye.

Ethers with an allyl position of the double bond. Part 7: Allyl  
ether of salicylaldehyde in the Grignard reaction. Zhur.ob.khim.  
30 no.10:3261-3263 0 '61. (MIRA 14:4)

1. Belorusskiy politekhnicheskiy institut.  
(Salicylaldehyde)

PANSEVICH-KOLYADA, V.I.

Oxidation of hydrocarbon having conjugated double bonds by organic per acids. Part 3: Mechanism of the oxidation of 1,3-diene hydrocarbons, unsymmetrically substituted by alkyl radicals, with acetyl hydroperoxide. Zhur. ob. khim. 30 no.12:3901-3907 D '60.  
(MIRA 13:12)

1. Belorusskiy politekhnicheskiy institut.  
(Olefins) (Oxidation) (Peroxyacetic acid)

PANSEVICH-KOLYADA, V.I.

Characteristics of the reaction of bromination of 1,1-diphenyl-2-bromo-1-propanol and 1,1-diphenyl-2-bromo-1-propene. Zhur. ob. khim. 30 no.12:3898-3901 D '60. (MIRA 13:12)

1. Belorusskiy politekhnicheskiy institut.  
(Propanol) (Propene) (Bromination)

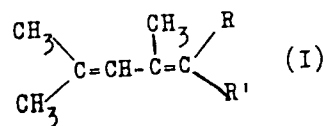
S/079/60/030/012/007/027  
B001/B064

AUTHOR: Pansevich-Kolyada, V. I.

TITLE: Oxidation of Hydrocarbons With Conjugate Double Bonds by Means of Organic Peracids. III. Oxidation Mechanism of 1,3-Diene Hydrocarbons With Unsymmetrically Substituted Alkyl Radicals by Means of Peracetic Acid

PERIODICAL: Zhurnal obshchey khimii, 1960, Vol. 30, No. 12, pp. 3901-3907

TEXT: The oxidation of the 1,3-diene hydrocarbons



either proceeds at a ratio of 1 mole hydrocarbon to 2 moles peracid on both double bonds, under the formation of dioxide (II) and monoacetates of oxydodiols with the possible structures (III)-(VIII), or at a molar ratio of 1:1 on one double bond under the formation of unsaturated mono-oxides with the possible structures (IX) or (X) and monoacetates of unsaturated diols of the possible structures (XI)-(XVI) (Refs. 1, 2). In the present paper the author continued to investigate the structure of the reaction products. The results obtained should be a basis for a better evaluation.

Card 1/4

Oxidation of Hydrocarbons With Conjugate Double Bonds by Means of Organic Peracids. III. S/079/60/030/012/007/027  
 Oxidation Mechanism of 1,3-Diene Hydrocarbons B001/B064  
 With Unsymmetrically Substituted Alkyl Radicals  
 by Means of Peracetic Acid

tion of the oxidation mechanism of the hydrocarbons with peracetic acid. The experimental results show that the oxidation of the 1,3-diene hydrocarbons, unsymmetrically substituted by alkyl radicals, proceeds with peracetic acid simultaneously into two directions, in 2,3 and 4,5 position under the formation of two corresponding, unsaturated monoxides. Monoxides which have the oxide ring in the 2,3 position, yield dioxides when further oxidized with peracetic acid. The monoxides with the 4,5 oxide ring are likely to acetylate (Refs. 22, 23) under the formation of monoacetates of unsaturated diols, which, on further oxidation with hydroperoxide, yield monoacetates of oxydodiols. The reactions mentioned follow, in contrast to the previous assumption of the author (Ref. 2) the principal reaction scheme:

Card 2/4



Oxidation of Hydrocarbons With Conjugate Double Bonds by Means of Organic Peracids. III. S/079/60/030/012/007/027  
Oxidation Mechanism of 1,3-Diene Hydrocarbons B001/B064  
With Unsymmetrically Substituted Alkyl Radicals  
by Means of Peracetic Acid

The structure of all oxidation products was determined. There are 25 references: 15 Soviet, 4 US, 4 French, and 2 German.

ASSOCIATION: Belorusskiy politekhnicheskiy institut  
(Belorussian Polytechnic Institute)

SUBMITTED: December 29, 1959

Card 4/4

SOV/79-29-6-32/72

5(3)  
 AUTHORS: Levina, S. A., Yermolenko, N. P., Pansevich-Kolyada, V. I.

TITLE: Ferric Hydroxide, a Catalyst of the Reaction of the Condensation of Acetone to Diacetone Alcohol (Gidrat oksid azolezatsii katalizator reaktsii kondensatsii acetona v diatsetonovyy spirt)

PERIODICAL: Zhurnal obshchey khimii, 1959, Vol 29, Nr 6, pp 1920 - 1925 (USSR)

ABSTRACT: The diacetone alcohol is obtained by condensation of acetone in the presence of catalysts. As such catalysts the hydroxides of the alkali metals (Refs 1-4), of calcium (Refs 5,6), barium (Ref 7), and some other products (Refs 8,9) are used. In the present paper the authors for this purpose used ferric hydroxide prepared in a suitable way (Ref 10). The synthesis was carried out according to the usual laboratory method (Ref 11). In the experiment many samples of ferric hydroxide were used, which were prepared by precipitation with ammonia from sulfate in different ways. The structure of the samples was determined, and their catalytic activity was compared with the activity of barium hydroxide. It was found that it was possible to obtain ferric hydroxide of different adsorption and catalytic

Card 1/2

Ferric Hydroxide, a Catalyst of the Reaction of the  
Condensation of Acetone to Diacetone Alcohol

SOV/79-29-6-32/72

activity according to the mode of preparation. The activity rises with decreasing content of the ion  $\text{SO}_4^{2-}$ . It was thus confirmed that the use of ferric hydroxide as catalyst for the condensation of acetone to the diacetone alcohol is possible. The constant of the condensation rate in the presence of the most active sample of ferric hydroxide (Sample II in table 1) is twice higher than in the presence of barium hydroxide (Tables and Figures). There are 4 figures, 3 tables, and 17 references, 6 of which are Soviet.

ASSOCIATION: Institut khimii Akademii nauk Belorusskoy SSR (Institute of Chemistry of the Academy of Sciences, Belorusskaya SSR)

SUBMITTED: March 3, 1958

Card 2/2

5(3)

SOV/79-29-6-33/72

AUTHOR: Pansevich-Kolyada, V. I.

TITLE: Investigation of the Ethers With Allyl Position of the Double Bond (Issledovaniye prostykh efirov s allil'nyim polozheniyem dvoynoy svyazi). V. The Formation of Ether of Furfuryl Alcohol With the  $\alpha, \alpha, \gamma, \gamma$ -Tetrasubstituted Allyl Alcohols (V. Eterifikatsiya furfurilovogo spirta  $\alpha, \alpha, \gamma, \gamma$ -tetrazameshchennymi allilovymi spirtami)

PERIODICAL: Zhurnal obshchey khimii, 1959, Vol 29, Nr 6, pp 1925 - 1928 (USSR)

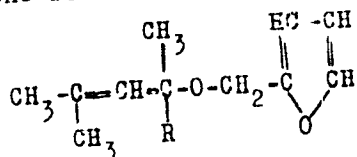
ABSTRACT: The number of ethers of the furfuryl alcohol is very limited. The easy cleavage of the furan cycle in acid medium, especially in heating, i.e. conditions under which the ethers are most easily obtained is an obstacle to their synthesis. The  $\alpha, \alpha, \gamma, \gamma$ -tetrasubstituted allyl alcohols - as was shown already earlier (Refs 10 - 12) - form ethers readily and in good yields. The formation of ether for these alcohols takes place only in primary alcohols of the aliphatic and the aromatic series, with glycol, glycerin, and other compounds having a primary alcohol group.

Card 1/3 In the present paper this characteristic feature of the tetra-

Investigation of the Ethers With Allyl Position of the Double Bond. V. The Formation of Ether of Furfuryl Alcohol With the  $\alpha, \alpha, \gamma, \gamma$ -Tetrasubstituted Allyl Alcohols

SOV/79-29-6-33/72

substituted alcohols was used for their production of ether by means of furfuryl alcohol. In the case of the action of a small amount of diluted sulfuric acid on the solutions of the 2,4-dimethyl pentene-2-ol-4 (I), 2,4-dimethyl hexene-2-ol-4 (II), 2,4-dimethyl heptene-2-ol-4 (III), 2,4-dimethyl octene-2-ol-4 (IV), 2,4,7-trimethyl octene-2-ol-4 (V) and 2-methyl-4-phenyl pentene-2-ol-4 (VI) in furfuryl alcohol (VII) ethers of furfuryl alcohol with the  $\alpha, \alpha, \gamma, \gamma$ -tetrasubstituted allyl alcohols of the following structure are formed:



(VIII)  $\text{R} = \text{CH}_3$ , (IX)  $\text{R} = \text{C}_2\text{H}_5$ , (X)  $\text{R} = n\text{-C}_3\text{H}_7$ , (XI)  $\text{R} = n\text{-C}_4\text{H}_9$ , (XII)  $\text{R} = i\text{-C}_5\text{H}_{11}$ , (XIII)  $\text{R} = \text{C}_6\text{H}_5$ . There are 1 table and 10 references, 14 of which are Soviet.

Card 2/3

Investigation of the Ethers With Allyl Position of the Double Bond. V. The Formation of Ether of Furfuryl Alcohol With the  $\alpha, \alpha, \gamma, \gamma$ -Tetrasubstituted Allyl Alcohols SOV/79-29-6-33/72

ASSOCIATION: Belorusskiy politekhnicheskii institut (Belorussian Polytechnic Institute)

SUBMITTED: May 22, 1958

Card 3/3

5(3)

SOV/79-29-4-34/77

AUTHORS:

Pansevich - Kolyada, V. I., Bogush, B. K.

TITLE:

Investigation of Esters With an Allyl Position of the Double Bond (Issledovaniye prostykh efirov s allil'nyy polozheniyem dvoynoy svyazi). IV. Preparation of Some Esters of 2-Methyl-4-benzyl-penten-2-ol-4 and 1-Phenyl-buten-1-ol-3 (IV. Polucheniye nekotorykh prostykh efirov 2-metil-4-benzil-penten-2-ola-4 i 1-fenilbuten-1-ola-3)

PERIODICAL:

Zhurnal obshchey khimii, 1959, Vol 29, Nr 4, pp 1198-1201 (USSR)

ABSTRACT:

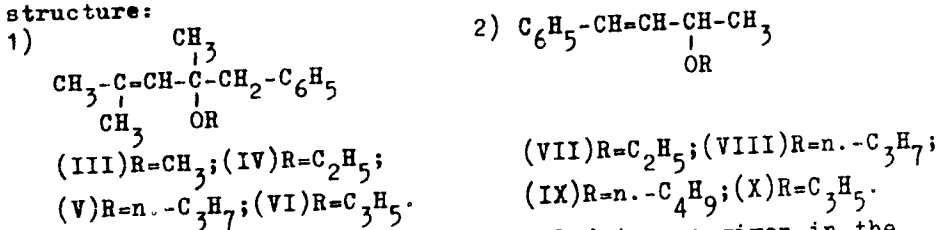
In the previous reports (Refs 1-3) published by the authors it was shown that  $\alpha,\alpha,\gamma,\gamma$ -tetra-alkyl-substituted allyl alcohols are exceptionally capable of forming esters with primary alcohols in the presence of a small quantity of diluted sulfuric acid. It is known that this formation of esters is also easily possible in the case of some tertiary alcohols containing a phenyl- or styryl radical (Refs 4-14). Therefore it seemed rather obvious to carry out the formation of esters of tetra-substituted allyl alcohols with aryl radicals. In the present article the results of the formation of ester of 2-methyl-4-benzyl-penten-2-ol-4 (I) and 1-phenyl-buten-1-ol-3 (II) with methyl-, ethyl-, n.-propyl-,

Card 1/3

SOV/79-29-4-34/77

Investigation of Esters With an Allyl Position of the Double Bond. IV. Preparation of Some Esters of 2-Methyl-4-benzyl-penten-2-ol-4 and 1-Phenyl-buten-1-ol-3

n.-butyl-, and allyl alcohol are described. It was found that the alcohol (I) is almost as easily transformed into the ester as the tetra-alkyl-substituted allyl alcohols (Refs 1-3), and that the ester formation of the alcohol (II) is much more difficult, as was noted on an earlier occasion (Ref 1). This different reaction rate in the ester formation of these two alcohols (I) and (II) may only be explained by the different nature of the alcohol groups. The synthesized esters exhibit the following structure:



The physical constants and analytical data are given in the table. Only benzoic acid aldehyde and benzoic acid were found in

Card 2/3

SOV/79-29-4-34/77

Investigation of Esters With an Allyl Position of the Double Bond. IV. Preparation of Some Esters of 2-Methyl-4-benzyl-penten-2-ol-4 and 1-Phenyl-buten-1-ol-3

the oxidation products of 1-phenyl-3-ethoxy-buten-1 with potassium permanganate. In all, eight **esters** hitherto unknown were synthesized, all of them with an allyl position of the double bond. There are 1 table and 17 references, 5 of which are Soviet.

ASSOCIATION: Belorusskiy politekhnicheskiy institut (Belorusskiy Polytechnic Institute)

SUBMITTED: March 6, 1958

Card 3/3

SOV/19-58-6-102/685

AUTHORS: Pansevich-Kolyada, V. I., and Osipenko, I. P.

TITLE: A Method of Obtaining 2,4-Dimethyl-4-Benzyl-Thiopentene-2.  
(Sposob polucheniya 2,4-dimetil-4-benziltiopentena-2)

PERIODICAL: Byulleton' izobreteniy, 1958, Nr 6, p 26 (USSR)

ABSTRACT: Class 12<sup>0</sup>, 19<sup>03</sup>. Nr 113780 (587368 of 6 Dec 1957).  
Submitted to the Committee for Inventions and Discoveries  
at the Ministers Council of USSR. A method as specified in  
title, differing from known methods by the following: 2,4-  
dimethyl-pentene-201-4 is mixed with benzyl-mercaptan con-  
taining sulfuric acid, and the formed thioether separated  
by known means.

Card 1/1

**AUTHORS:** Pansevich-Kolyada, V. I., Osipenko, I. F. 79-28-4-13/60

**TITLE:** The Investigation of Simple Ethers With an Allyl Position of the Double Binding (Issledovaniye prostykh efirov s allil'nym polozheniyem dvoynoy svyazi).  
III. An Investigation of the Etherification Reaction of  $\alpha, \alpha, \gamma, \gamma$ -Tetraalkylsubstituted Allyl Alcohols (III. Izucheniye reaktsii eterifikatsii  $\alpha, \alpha, \gamma, \gamma$ -tetraalkilzameshchennykh allilovykh spirtov)

**PERIODICAL:** Zhurnal Obshchey Khimii, 1958, Vol. 28, Nr 4, pp. 909-914 (USSR)

**ABSTRACT:** In the present paper the authors investigated the influence of the structure of primary alcohols on the etherification products of tetraalkyl substituted allyl alcohols, the nature of the alcohol group (primary, secondary, tertiary), the possibility of etherification with heterofunctional compounds containing a primary alcohol group, as well as the possibility of an ether formation of the investigated allyl alcohols with 2- and 3-atomic alcohols. For this purpose an etherification was conducted

Card 1/3

The Investigation of Simple Ethers With an Allyl  
Position of the Double Binding. III. An Investigation  
of the Etherification Reaction of  $\alpha$ ,  $\alpha$ ,  $\gamma$ ,  $\gamma$ -Tetraalkyl-  
substituted Allyl Alcohols

79-28-4-13/60

ted with various hydroxyl containing compounds, 2,4 dimethylpentene -2-ol-4 (I), 2,4-dimethylhexene-2-ol-4 (II), 2,4-dimethylheptene-2-ol-4 (III) and dimethyloctene -2-ol-4 (IV). Allyl ethers of polyoxy compounds have been described repeatedly in publications (References 3-10). Some of these possess the valuable property of forming durable transparent and insoluble films (References 3,4). This induced the authors to investigate the etherification reaction of the  $\alpha$ ,  $\alpha$ ,  $\gamma$ ,  $\gamma$ -tetraalkyl substituted allyl alcohols with 2- and 3-atomic alcohols. The reaction of the etherification of the alcohols (II, III, IV) with ethyleneglycol takes an analogous course to that with primary monoatomic alcohols and leads to the simultaneous formation of the respective mono- and diethers of ethylene glycol (XII-XVI). With glycerin glycerinediethers are formed. The influence of the hydrocarbon radical of the primary monoatomic alcohol on the reaction velocity was investigated in the example of the etherification of the alcohol (IV) with equimolecular amounts of mixed methyl- and n-

Card 2/3

The Investigation<sup>of</sup> Simple -Ethers With an Allyl 79-28-4-13/60  
Position of the Double Binding. III. An Investigation  
of the Etherification Reaction of  $\alpha$ ,  $\alpha$ ,  $\gamma$ ,  $\gamma$ -Tetraalkyl-  
substituted Allyl Alcohols

butyl alcohol. With methyl alcohol the yield was twice  
that with n-butyl alcohol. Therefore the reaction velo-  
city depends upon the size of the hydrocarbon radical  
of the primary alcohol.  
There are 1 table and 12 references, 7 of which are  
Soviet.

ASSOCIATION: Belorusskiy lesotekhnicheskiy institut (Belorussian  
Institute for Forestry)

SUBMITTED: April 15, 1957

Card 3/3

**AUTHORS:** Pansevich-Kolyada, V. I., Idel'chik, Z. B. 79-28-4-14/60

**TITLE:** Investigations in the Field of Oxide Compounds (Issledovaniya v oblasti oksidosoyedineniy).  
 XI. The Interaction of  $\alpha$ -Oxides of the Allyl Ethers of Phenol, p-Cresol and Guaiacol With Methyl Alcohol  
 (Vzaimodeystviye  $\alpha$ -okisey allilovykh efirov fenola, p-krezola i gvayakola s metilovym spirtom)

**PERIODICAL:** Zhurnal Obshchey Khimii, 1958, Vol. 28, Nr 4, pp. 914-916 (USSR)

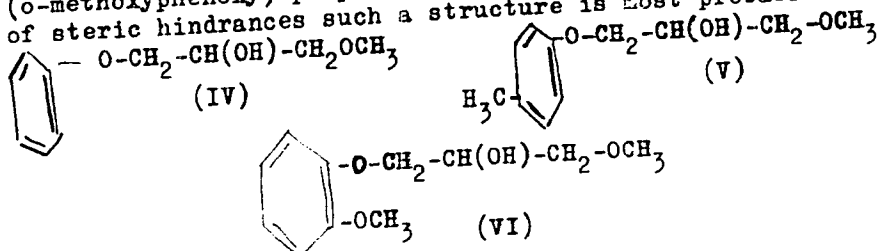
**ABSTRACT:** Usually, incomplete glycol ethers are formed as reaction products of the interaction of  $\alpha$ -oxides with alcohols. The structure of these ethers is determined by the sequence of the oxide nucleus fission, which again is dependent upon the structure of the  $\alpha$ -oxides and on the character of the used catalyst. In the present paper the authors investigated the interaction reaction of the  $\alpha$ -oxides of phenol - (I) of p-cresol - (II) and of guaiacol - (III) allyl ethers with methyl alcohol, in the presence

Card 1/3

79-28-4-14/60

Investigations in the Field of Oxide Compounds.  
 XI. The Interaction of  $\alpha$ -Oxides of the Allyl Ethers  
 of Phenol, p-Cresol and Guaiacol With Methyl Alcohol

of sodium methylate. They found the presence of mixed incomplete glycerin ethers among the reaction products. The fission of the three-membered oxide cycle takes place after the binding of the oxygen oxide with a more strongly hydrated carbon atom in the presence of alkaline catalysts, as was found in a series of the  $\alpha$ -oxides. For this reason the glycerin ethers obtained by the authors presumably have the structure: 1-methoxy-3 phenoxy-propanol-2 (IV), 1-methoxy-3-(p-tolyloxy)-propanol-2 (V), 1-methoxy-3-(o-methoxyphenoxy)-propanol-2 (VI). From the viewpoint of steric hindrances such a structure is most probable.



Card 2/3

Investigations in the Field of Oxide Compounds.

79-28-4-14/60

XI. The Interaction of  $\alpha$ -Oxides of the Allyl Ethers  
of Phenol, p-Cresol and Guaiacol With Methyl Alcohol

There are 21 references, 15 of which are Soviet.

ASSOCIATION: Institut khimii Akademii nauk Belorusskoy SSR i Belorusskiy politekhnicheskiy institut Belorussiya, (Institute for Chemistry, AS Belorussian SSR, and Belorussian Polytechnical Institute)

SUBMITTED: March 28, 1957

Card 3/3

79-28 3-17/61

AUTHORS:

Pansevich-Kolyada, V. I.,  
Osipenko, I. F.

TITLE:

Investigation of the Ethers With an Allyl Position of  
the Double Bond (Issledovaniye prostykh efirov s  
allil'nyim polozheniyem dvoynoy svyazi).  
II. Synthesis of the Ethers of  $\alpha, \alpha, \beta, \beta$ -Tetraalkyl-  
Substituting Allyl Alcohols (Sintez prostykh efirov  $\alpha, \alpha, \beta, \beta$ -  
tetraalkilzameshchennykh allilovykh spirtov)

PERIODICAL:

Zhurnal Obshchey Khimii 1958, Vol. 28, Nr 3, pp. 641-646  
(USSR)

ABSTRACT:

Aryl-substituted tertiary alcohols easily form ether  
(ref. 1) in the presence of some acids and hydrochlorides  
and still more easily aryl-substituted alcohols having  
a double bond in the allyl position (refs. 2-5). Following  
an earlier work the etherization of 2,4-dimethylpentene-  
2-ols-4- (formula I); 2,4-dimethylhexene-2-ols-4- (II);  
2,4-dimethylheptene-2-ols-4 (III) and 2,2-dimethyloctene-  
2-ols-4 (IV) with primary monovalent alcohols of the fatty  
series was worked out by the authors. The ethers obtained here

Card 1/3

Investigation of the Ethers With an Allyl Position of the <sup>79</sup>-28-3-17/61  
Double Bond.

II. Synthesis of the Ethers of  $\alpha, \alpha, \beta, \gamma$ -Tetraalkyl-Substituting Allyl  
Alcohols

had the structure of the formulae (V) to (XVI). The formation of all mentioned ethers takes place very easily and mostly exothermally. The yields in the etherization with methylalcohol are the greatest and become smaller with the use of higher monovalent alcohols, their structure practically exercising no influence on them. In order to determine the influence of the structure of tertiary alcohols on the etherization the authors tried to convert the 2-methylbutanol - 2 (XVII), 2-methylbutene-3 - ol - 2 (XVIII), 3-methylhexene - 5 - ol - 3 (XIX) and 2,4-dimethylexanol-4 (XX) to ether. These compounds as well as (I to IV) are tertiary alcohols, but different from these they are saturated. The etherization of the alcohols (XVII-XX) with methyl- and ethylalcohol was not successful. The usual easy ether formation of alcohols (I) to (IV) is based on their structure. The tertiary alcohols are more easily converted to alcohol than secondary and primary ones as their accumulation of radicals on the carbinolhydrocarbon

Card 2/3

Investigation of the Ethers With an Allyl Position of the <sup>79</sup>-28 3-17/61  
Double Bond.

II. Synthesis of the Ethers of  $\alpha, \alpha, \gamma, \gamma$  -Tetraalkyl-Substituting Allyl  
Alcohols

atom leads to a weakening of the C - O - bond. It was found that the presence of the vinyl radical semisubstituted by methyl groups on the carbinol hydrocarbon atom is the reason for the easy etherization. The ether synthesis of tetra-alkylsubstituted allyl alcohols takes place under substitution of the hydroxyl of tertiary alcohol by the alkoxygroup.

There are 2 tables and 16 references, 3 of which are Soviet

ASSOCIATION: Belorusskiy lesotekhnicheskiy institut  
(Belorussian Institute for Forestry Engineering)

SUBMITTED: March 14, 1957

Card 3/3

**AUTHORS:** Idel'chik, Z. B., Pansevich-Kolyada, V. I. 79-28 3-02, 01

**TITLE:** Investigations Within the Field of Oxide Compounds (Issledovaniya v oblasti oksidosoyedineniy). X. The Reaction of the  $\alpha$ -Oxides of the Allylethers of Phenol, o- and p-Cresols and Guaiacol With Amines (X. Vzaimodeystviye  $\alpha$ -okisey allilovykh efirov fenola, o-, p-krezolov i gvayakola s aminami)

**PERIODICAL:** Zhurnal Obshchey Khimii, 1958, Vol 28, Nr 3, pp. 792-795 (USSR)

**ABSTRACT:** Some derivatives of the  $\alpha$ -oxides of the allylethers of phenols are, as is known, physiologically active preparations (references 1,2). Of the same effect are also some derivatives of the  $\alpha$ -oxides of the allylethers of nitrophenols (Ref. 3). In this work the authors aimed at deepening the knowledge on such compounds by investigating the reactions of the  $\alpha$ -oxides of the allylethers of phenol, o- and p-cresols and guaiacol with diethylamine, as well as the reactions of the  $\alpha$ -oxides of the allylethers of phenol and o-cresol with piperidine. 1-phenoxy-2,3-propane oxide (I), 1-(o-tolyloxy)-2,3-oxydopropane (II), 1-(p-tolyloxy)-2,3-oxydopropane (III), and 1-(o-methoxyphenoxy)-2,3-oxydopropane (IV) - all these oxides of allylphenoether were synthesized by the

Card 1/3

79-28-5 11/6  
 .Investigations Within the Field of Oxide Compounds. X. The Reaction of the  $\alpha$ -Oxides of the Allylethers of Phenol, o- and p Cresols and Guaiacol With Amines

oxidation of the allylethers of the corresponding phenols with acetylhydrogenperoxide ( $\text{CH}_3\text{CO} - \text{O} - \text{OH}$ ) as well as by the reaction of the phenols with epichlorohydrine. By slight heating of these etheroxides (I, II, III and IV) with diethylamine as well as of the ether oxides with piperidine the corresponding phenoxyaminoalcohols were obtained. Based on known papers (References 1,2) on the reaction of the glycidic ethers of the phenols with amines the structure of the synthesized phenoxyaminoalcohols can be expressed by the formulae: 1: phenoxy-3-diethylaminopropanol-2 (V), 1-(o-tolyloxy)-3-diethylaminopropanol-2 (VI), 1-(p-tolyloxy)-3-diethylaminopropanol-2 (VII), 1-(o-methoxyphenoxy)-3-diethylaminopropanol-2 (VIII), 1-phenoxy-3-(N-piperidyl)-propanol-2 (IX), and 1-(o-tolyloxy)-3-(N-piperidyl)-propanol-2 (X). All these alcohols easily form hydrochlorides and iodoethylates in water. The compounds (VI, IX and X) are only produced in form of their derivatives. The pharmacological investigation of the phenoxyaminoalcohols (V-X) and of their hydrochlorides and iodoethylates showed that they have an high hypotensive effect with low toxic effect. The toxic effect increases from the free alcohols to the hydrochloride

Card 2/3

Investigations Within the Field of Oxide Compounds. X. The Reaction <sup>79-28-3-1161</sup>  
of the  $\alpha$ -Oxides of the Allylethers of Phenol, o- and p-Cresols and Guaiacol  
With Amines

salts and iodoethylates, the hypotensive effect, however, of  
these preparations increasing simultaneously, and at the same  
ratio. There are 10 references, 3 of which are Soviet.

ASSOCIATION: Institut khimii Akademii nauk Belorusskoy SSR (Chemical  
Institute, AS Belorussian SSR)

SUBMITTED: March 28, 1957

Card 3/3

PANSEVICH-KLOYADA, V.I.; OSIPENKO, I.F.

Study of ethers with an allyl positioned double bond. Part 2:  
Synthesis of ethers of  $\alpha, \alpha, \gamma, \gamma$  -tetraalkylsubstituted allyl  
alcohols. Zhur. ob. khim. 28 no.3:641-646 Mr '58. (MIRA 11:5)

1. Belorusskiy lesotekhnicheskiy institut.  
(Ether)

PANSEVICH-KOLYADA, V.I.

2-phenyl-3-bromo-2-butene in Grignard reaction. Zhur.ob.khim. 28  
no.2:438-441 P '58. (MIRA 11:4)

1.Belorusskiy politekhnicheskiy institut.  
(Butene) (Grignard reagents)

PANSEVICH-KOLYADA, V.I.; OSIPIYKO, I.F.

Study of ethers with allyl position of double bond. Part 3: Study  
of  $\alpha, \alpha, \beta, \beta$ -tetraalkyl substituted allyl alcohol etherification  
reaction. Zhur. ob. khim. 28 no.4:909-914 Ap '58. (MIRA 11:5)

1. Belorusskiy lesotekhnicheskiy institut.  
(Allyl alcohol) (Etherification)

PANSEVICH-KOLYADA, V.I.; IDEL'CHIK, Z.B.

Study of oxides. Part 10: Interaction of  $\alpha$ -oxides of allyl ethers of phenol,  $\beta$ -cresol, and guaiacol with methyl alcohol. Zhur. ob. (MIRA 11:5)  
khim. 28 no. 4:914-916 Apr '58.

1. Institut khimii Akademii nauk Belorusskoy SSR i Belorusskiy politekhnicheskii institut.  
(Oxides) (Ethers) (Alcohols)

PANSEVICH-KOLIADA, V.I.

Oxidation of alkadiene hydrocarbons of nonsymmetrical structure  
with acetyl peroxide. Part 2. Zhur. ob. khim. 26 no.8:2161-2170  
Ag '56. (MLBA 10:11)

1. Belorusskiy politekhnicheskiy institut.  
(Olefins) (Acetyl peroxide)

PANSEVICH-KOLYADA, V.I.

AZANOVSKAYA, M.M.; PANSEVICH-KOLYADA, V.I.

$\alpha$ -oxides of alkylfuryl alcohols. Zhur. ob. khim. 27 no.2:384-387 P '57. (MLRA 10:6)

1. Institut khimii Akademii nauk Belorusskoy SSR.  
(Furfuryl alcohol)

PANSEVICH-KOLYADA, V. I.

CHIZHEVSKAYA, I.I.; PANSEVICH-KOLYADA, V.I.

Research in the field of oxide compounds. Part 8: Reaction of allyl ether  $\alpha$ -oxides of o-, m-, and p-nitrophenols with diethylamine. Zhur.ob.khim. 27 no.5:1223-1226 My '57. (MLRA 10:8)

1. Institut khimii Akademii nauk Belorusskiy SSR.  
(Phenol) (Diethylamine)

CHIZHEVSKAYA, I.I.; PANGSVICH-KOLYADA, V.I.

Investigations in the field of oxide compounds. Part 9: Reaction of  $\alpha$ -oxides of allyl ethers of nitrophenols with benzimidazole. Zhur.ob. khim. 27 no.6:1495-1498 Je '57. (LRA 10:8)

1. Institut khimii Akademii nauk Belorusskoy SSR.  
(phenols) (Benzimidazole) (Propanol)

PANSHIN, A., general-mayor

Military builders undergo a difficult test. Komm.Vooruzh.Sil 1  
no.17:46-51 S '61. (MIRA 14:8)  
(Military engineers)

BEIYAK, Yu.L., kand. tekhn. nauk; PANSIN, A.F., inzh.

Effect of wave impact on ship hulls. Sudostroenie 25 no.7:10-11  
Jl '59. (MIRA 12:12)  
(Ships--Hydrodynamic impact) (Hulls (Naval architecture))

PANSHIN, A.P.

Modernizing loose-bulk cargo deck barges with a 3,000-ton loading capacity. Rech.transp. 18 no.12:26-27 D '59.

(MIRA 13:4)

(Barges)

PANSHIN, A.F., inzh.

Considering cushion rigidity in determining the frequency of  
free vibrations in large tonnage heavy boat hulls. Trudy  
TSNIIRF no.40:69-74 '59. (MIRA 13:6)  
(Vibration (Marine engineering))  
(Hulls (Naval architecture))

PANSHIN, A. F.

Name: PANSHIN, A. F.

Dissertation: Investigation of the dynamic stability under wave conditions  
of large-tonnage vessels for inland transportation

Degree: Cand Tech Sci

~~DEFENSE AT~~

~~Association:~~ Min Higher Education USSR, Gor'kiy Polytechnical Inst imeni  
A. A. Zhdanov

~~PUBLICATION~~

~~Defense Date,~~ Place: 1956, Gor'kiy

Source: Knizhnaya Letopis', No 52, 1956

PANSHIN, A.F., insh.

Some remarks on the method proposed by V.D.Drinkov for calculating  
the strength of ships used in inland navigation. Rech. transp. 17  
no.8:55-3 of cover Ag '58. (MIRA 11:10)  
(Hulls--Testing)

PANSHIN, A.F.

Treatment of onychomycosis. Vest. vener. No.3:50-51 May-June 50.  
(CLML 19:4)

1. Of the Clinic for Skin and Venereal Diseases (Head -- Prof.  
S.Ya. Golosovker) Leningrad State Pediatric Medical Institute.

PANSHIN, A.P., inch.

Importation of the "Vichard" type motor ships  
by the USSR. (LVT No. 14:58-64 '61.  
(REF 14:11)

(L. Dept., Indo-Ship trials)

SOV/124-57-3-3095

Translation from: Referativnyy zhurnal. Mekhanika, 1957, Nr 3, p 63 (USSR)

AUTHOR: Belyak, Yu. L., Panshin, A. F.

TITLE: Experimental Determination of the Wave Stresses Exerted on the Hull of Oil Tankers (Eksperimental'noye opredeleniye volnovykh nagruzok na korpusa neftenalivnykh sudov)

PERIODICAL: Tr. Tsentr. n-i. in-ta rech. flota, 1956, Nr 32, pp 16-42

ABSTRACT: Results are given of full-scale strength tests on three types of oil barges of different tonnage. Bending moments, developing stresses of the order of the main allowable stresses, were created by a suitable distribution of ballast. This enabled the authors to conduct a verification of the hull strength under the conditions indicated. Seaway tests of the vessels were conducted off the Astrakhan' roadstead with the aim of determining the stresses in the hull joints and the character of hull flexure in a seaway. The tests revealed that maximum additional stresses created by the action of waves upon the hull develop when the vessel is proceeding on a quartering or "off-the-bow" course. Propositions are tendered for changes of the "Instruction for the verification of the strength of vessels"

Card 1/2